

AD-A088 043

CALIFORNIA UNIV SANTA BARBARA

F/G 7/4

SULFUR MODIFIES THE CHEMISORPTION OF CARBON MONOXIDE ON RHODIUM--ETC(U)

JUL 80 R M KROEGER, W C KASKA, P K HANSHA

N00014-78-C-0011

UNCLASSIFIED

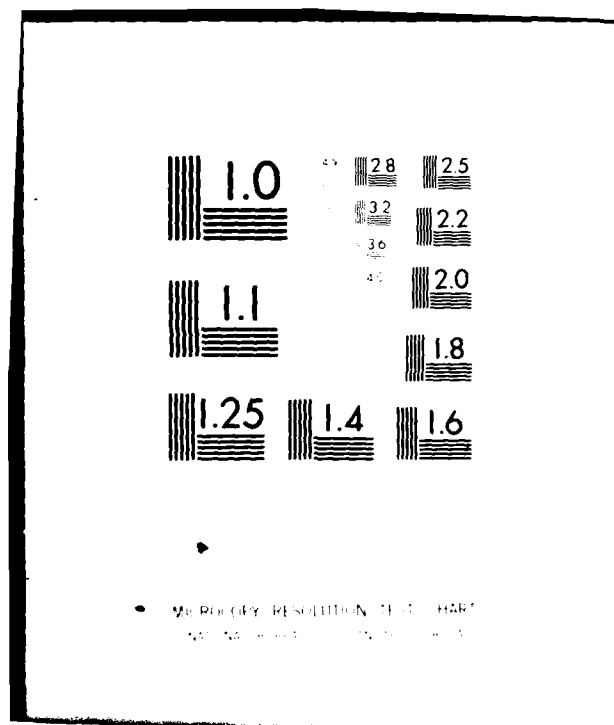
TR-3

NL

1-1
AU
AD-A088 043



END
DATE
FILMED
DTIC



AD A088043

LEVEL II

12

OFFICE OF NAVAL RESEARCH

Contract N00014-78-C-0011

Task No. NR 056-673

TECHNICAL REPORT NO. 3✓

Sulfur Modifies the Chemisorption
of Carbon Monoxide on Rhodium/Alumina
Model Catalysts

by

R.M. Kroeker, W.C. Kaska and P.K. Hansma

Prepared for Publication

in the

Journal of Catalysis

University of California, Santa Barbara
Department of Chemistry
Department of Physics
Quantum Institute
Santa Barbara, CA

July 1980

DTIC
ELECTE
AUG 20 1980
S D C

Reproduction in whole or in part is permitted for
any purpose of the United States Government.

This document has been approved for public release
and sale; its distribution is unlimited.

DDC FILE COPY

80 8 19 014

UNCLASSIFIED

SECURITY CLASSIFICATION OF THIS PAGE (When Data Entered)

REPORT DOCUMENTATION PAGE		READ INSTRUCTIONS BEFORE COMPLETING FORM
1. REPORT NUMBER	2. GOVT ACCESSION NO.	3. RECIPIENT'S CATALOG NUMBER
	AD-A088043	
4. TITLE (and Subtitle)		5. TYPE OF REPORT & PERIOD COVERED
SULFUR MODIFIES THE CHEMISORPTION OF CARBON MONOXIDE ON RHODIUM/ALUMINA MODEL CATALYSTS,		(9) Technical Report
		6. PERFORMING ORG. REPORT NUMBER
7. AUTHOR(s)		8. CONTRACT OR GRANT NUMBER(s)
R.M./Kroeker, W.C./Kaska W.C. P.K./Hansma		(15) N00014-78-C-0011
9. PERFORMING ORGANIZATION NAME AND ADDRESS		10. PROGRAM ELEMENT, PROJECT, TASK AREA & WORK UNIT NUMBERS
Department of Chemistry, Department of Physics & Quantum Institute, Univ. of Calif., Santa Barbara, CA 93106		NR 056-673 (1)
11. CONTROLLING OFFICE NAME AND ADDRESS		12. REPORT DATE
Office of Naval Research Department of the Navy Arlington, VA 22217 (11) 31 Jul		7/29/80
14. MONITORING AGENCY NAME & ADDRESS (if different from Controlling Office)		13. NUMBER OF PAGES
(14) TR-3		13
		15. SECURITY CLASS. (of this report)
		UNCLASSIFIED
		15a. DECLASSIFICATION/DOWNGRADING SCHEDULE
16. DISTRIBUTION STATEMENT (of this Report)		
Approved for public release and sale; distribution unlimited.		
17. DISTRIBUTION STATEMENT (of the abstract entered in Block 20, if different from Report)		
18. SUPPLEMENTARY NOTES		
19. KEY WORDS (Continue on reverse side if necessary and identify by block number)		
Sulfur, carbon monoxide, rhodium, poisoning, tunneling spectroscopy, vibrational spectroscopy		
20. ABSTRACT (Continue on reverse side if necessary and identify by block number)		
<p>The effect of D₂S on the chemisorption of CO on supported rhodium is examined with tunneling spectroscopy. It is seen that for low concentrations, D₂S decreases the density of species with two CO molecules per Rh atom and increases the density of species with one CO molecule per Rh atom. At higher concentrations D₂S decreases the density of all types of adsorbed CO.</p>		

DD FORM 1 JAN 73 1473

EDITION OF 1 NOV 65 IS OBSOLETE

UNCLASSIFIED 07.45
SECURITY CLASSIFICATION OF THIS PAGE (When Data Entered)

Sulfur Modifies the Chemisorption
of Carbon Monoxide on Rhodium/Alumina
Model Catalysts

R. M. Kroeker, W. C. Kaska, and P. K. Hansma*

Department of Physics (PK and PH) and

Department of Chemistry (WK)

University of California, Santa Barbara, California 93106

ABSTRACT

The effect of D_2S on the chemisorption of CO on supported rhodium is examined with tunneling spectroscopy. It is seen that for low concentrations, D_2S decreases the density of species with two CO molecules per Rh atom and increases the density of species with one CO molecule per Rh atom. At higher concentrations D_2S decreases the density of all types of adsorbed CO.

*Alfred P. Sloan Foundation Fellow

Accession For	
NTIS GRA&I	
EDC TAB	
Unannounced	
Justification	
By	
Distribution/	
Availability Codes	
Dist	Avail and/or special
A	

I. INTRODUCTION

There is considerable interest in the effect of sulfur on supported metal catalysts because of industrial catalyst poisoning by sulfur compounds during the hydrogenation of coal and processing of petroleum raw materials.¹ The nature of sulfur bonding to supported transition metals is not, in general, known. There are, however, model systems of small metal clusters which contain sulfur.² These molecular systems can form the basis of understanding some of the reactions of sulfur compounds with supported metals.³

There has been infrared work on the effect of sulfur on carbon monoxide adsorption. Recently Rewick and Wise⁴ observed that, in the presence of hydrogen, H_2S would preferentially desorb β -CO (multisite attached) over α -CO (single site attached) on alumina-supported Ni. In fact, under some conditions the amount of α -CO would actually increase for small H_2S exposure. At larger exposures it would decrease, but never as readily as β -CO. This type of measurement is exciting because it may lead to an understanding of selective catalyst poisoning.

We have chosen to study the effect of sulfur on carbon monoxide adsorption on alumina-supported Rh because: 1) the system of carbon monoxide on alumina-supported Rh has been extensively studied^{5,6} and is well understood, 2) there are three different types of adsorbed CO species to observe, and 3) the three types are agreed upon by both infrared and tunneling workers and can be distinguished from each other by either spectroscopy.

II. EXPERIMENTAL METHODS

The details of the preparation of inelastic electron tunneling spectroscopy tunnel junctions have been reported elsewhere.⁷ A brief description of the procedure follows:

Aluminum electrodes are deposited in high vacuum onto a glass slide. These thin aluminum films, usually 800 Å thick, are oxidized at 200°C in air to form the alumina insulating barrier necessary for tunneling. The oxidized electrodes are cleaned in an argon glow discharge prior to the deposition of rhodium. The rhodium is evaporated either in the presence of a CO plus D₂S gas mixture or in high vacuum, followed by exposure to the gas mixture. A gas mixture pressure of 1.3×10^{-3} Pascal (10^{-5} torr) is sufficient to dominate the residual gases present, typically 1.3×10^{-5} Pascal (10^{-7} torr) H₂O and H₂, and is maintained in the system for 700 sec to achieve a total exposure of order 1 Pascal-sec. Junctions are then completed by the evaporation of the top lead electrode. In this work all doped junctions were formed adjacent to undoped, control junctions to allow the subtraction of background structure. The differential spectrometer used has been described in the literature.⁸

III. EXPERIMENTAL RESULTS AND DISCUSSION

With no added sulfur, CO chemisorbs in at least three different ways on alumina supported rhodium particles.⁶ There are two different linear-bonded species (one CO per rhodium atom and two CO's per rhodium atom) and at least one multiply bonded species. The two linear-bonded species can be distinguished with tunneling spectroscopy by their different low energy modes. Specifically, the one with one CO molecule per

rhodium atom has a bending vibration at 58.1 meV while the one with two CO molecules per rhodium atom has a bending vibration at 51.6 meV⁶ (multiply by 8.065 cm⁻¹/meV to convert to cm⁻¹). The rhodium carbon stretching vibrations for these linear species appear as broad overlapping bands of low intensity centered at 73 meV and 70 meV respectively.

The relative amounts of the different species can be changed. For example, in the limit of "ultradispersion"⁹ all of the Rh atoms can adsorb two CO molecules. We can model an ultradispersed catalyst by simply evaporating a very small amount of Rh onto our oxidized Al (typically 1/4 monolayer average coverage as opposed to the one monolayer averaged coverage reported on previously⁶). Fig. 1 shows the low energy region for such a model catalyst. Note that for pure CO (0% D₂S) there is only one resolved peak, at 51.6 meV. This is the bending mode assigned to two CO molecules per Rh atom.

As sulfur is added, this peak decreases in intensity and a peak near 60 meV grows. This position is close to the 58.1 meV peak position for one CO molecule per Rh atom.⁶ It clearly suggests that the sulfur is displacing one of the CO molecules and leaving the other. The rhodium carbon stretching modes are seen only as broad shoulders of low intensity in Fig. 1, and are thus difficult to interpret. An additional mode of low intensity near 42 meV appears after the rhodium is exposed to D₂S. It is tentatively assigned as a rhodium-sulfur vibration. Quantitative calibration of the observed d²V/dI² spectra has not been attempted because of the difficulty (at present) of producing identical rhodium coverages.

In preliminary experiments with H_2S some characteristic hydrocarbon peaks were present in the spectra. To determine whether the source of the hydrogen in the hydrocarbons was the H_2S , we performed all subsequent experiments with D_2S . We never saw any CD stretching vibrations and concluded that the source of the hydrocarbon peaks was impurities. We have not observed any other modes possibly due to deuterium from the D_2S . Either they are not present or are of low intensity. With a mass spectrometer, we saw a family of deuterium containing fragments including D_2 , HD, D_2O , and HDO in the gas phase. We do not know which are liberated from the junction and which are due to exchange on the walls and filaments in the chamber.

We also performed some experiments with D_2S or H_2S preadsorption. We found that relatively low exposures ($\approx 3 \times 10^{-2}$ Pascal sec) would totally poison the catalyst for CO adsorption. In contrast, a larger exposure (10^{-1} Pascal sec) after CO adsorption would not totally desorb the CO. It would only decrease the peaks due to two CO molecules per Rh atom and the multiply bonded CO to roughly 1/10 of their original value. The peak due to one CO molecule per Rh atom remained the same or even increased. Its position remained near 58 meV. These experiments suggest that we are not observing sulfur-induced adsorption of CO but rather CO replacement by the D_2S .

We also prepared some junctions with no sulfur exposure, measured their spectra, and then heated the completed junctions in roughly 10^5 Pascal (1 atm.) of H_2S . At room temperature we observed a negligible effect for exposures as high as 10^8 Pascal sec. At 110°C an exposure of roughly 10^7 Pascal sec gave spectra comparable to those obtained by exposing uncompleted junctions (without the top lead electrode) to 10^{-1}

Pascal sec. Heating to 110°C without H₂S exposure does not give comparable spectra. Thus H₂S can diffuse through the top electrode, but the top electrode provides a significant barrier to this diffusion. The diffusion of a number of other substances through the top electrode has recently been studied by Jaklevic and Gaerttner.¹⁰ These experiments also suggest that we are observing CO replacement by D₂S and not sulfur-induced adsorption, because in this case only D₂S was present in the diffusion chamber.

Finally, we performed some experiments with very low CO exposures to determine if a species formed by adding just one CO molecule to a site that could absorb two CO molecules would resemble the species formed at CO saturation on the sites that could absorb only one CO molecule. Figure 2 shows the low energy region for a model catalyst with roughly 1 monolayer averaged Rh coverage (not ultradispersed). At the lowest CO exposures a peak near 58 meV dominates. By 3L (1.3×10^{-4} Pascal-sec) exposure the spectrum resembles a saturation spectrum⁶ with larger intensity in a 51.2 meV peak. Thus a peak at the position of one CO molecule per rhodium atom dominates for very low exposure while at higher exposures a peak at the position of two CO molecules per rhodium atom is larger. A reasonable interpretation is that at the low exposures (1L) some rhodium atoms that could accommodate two CO molecules have only had a chance to chemisorb one. This implies that the species formed by adding just one CO molecule to a site that could

adsorb two has a very similar bending vibrational frequency to that formed by adding one CO molecule to a site that can adsorb only one. The broad rhodium-carbon stretching mode for the single CO per rhodium atom species is seen at 73 meV in the 1L exposure spectrum. The addition of the metal-carbon stretching mode due to the two CO's per rhodium atom species at 70 meV cannot be resolved in the 3L exposure spectrum.

IV. SUMMARY

We have observed that:

- 1) the species with one CO molecule per Rh atom is displaced less readily by sulfur than the species with two CO molecules per Rh atom or the multiply bonded species;
- 2) the species with two CO molecules per Rh atom will convert to a species with one CO molecule per Rh atom in the presence of sulfur. This new species is similar, but distinguishable from the species with one CO molecule per Rh atom that forms in the absence of sulfur;
- 3) hydrogen sulfide will diffuse through the top electrode to modify the adsorbed CO on Rh. The resulting spectrum is similar to that obtained by posttreatment with hydrogen sulfide before the top electrode is evaporated. However, higher exposures and temperatures are required.

V. ACKNOWLEDGEMENTS

We want to thank Professor Gabor Somorjai for suggesting that a study of sulfur poisoning would be important and accessible by our technique. We want to thank Drs. John Yates, Dave Yates, and Jim Schwartz for helpful discussions and

L. Dubois for a critical reading of the manuscript and useful suggestions for improving it. This work was supported in part by the Office of Naval Research and the Materials Research Division of the National Science Foundation.

References

1. D. D. Whitehurst, ACS Symposium Series No. 71, "Organic Chemistry of Coal," edited by J. W. Larson.
2. a) H. Vahrenkamp, Angew. Chem. Internat. Edit. 14, 322 (1975).
b) J. L. Vidal, R. A. Fiato, L. A. Cosby and R. L. Pruett, Inorg. Chem. 17, 2574 (1978).
3. Some examples are:
 - a) R. J. Madon and H. Shaw, Catalysis Rev. 15, 69 (1977), and references therein.
 - b) P. C. H. Mitchell, Reactions on Sulfide Catalysts, in Catalysis Vol. I, Specialist Periodical Reports, The Chemical Society, 1977, p. 204.
 - c) G. A. Somorjai, Catalysis Rev. 18, 173 (1978).
 - d) T. E. Madey, J. T. Yates, Jr., D. R. Sandstrom and R. J. H. Voorhoeve, M. W. Roberts, and C. S. McKee, Chemistry of the Metal-Gas Interface, Clarendon Press, Oxford 1978, p. 542.
4. R. T. Rewick and H. Wise, J. Phys. Chem. 82, 751 (1978).
5. Some IR spectroscopy studies:
 - a) A. C. Yang and C. W. Garland, J. Phys. Chem. 61, 1504 (1957).
 - b) H. Arai and H. Tomanaga, J. Catalysis 43, 131 (1976).
 - c) H. C. Yao and W. G. Rothschild, J. Chem. Phys. 68, 4774 (1978).
 - d) J. T. Yates, T. M. Duncan, S. D. Worley and R. W. Vaughan, J. Chem. Phys. 70, 1219 (1979).
6. Tunneling spectroscopy studies:
 - a) R. M. Kroeker, W. C. Kaska and P. K. Hansma, J. Catal. 57, 72 (1979).
 - b) P. K. Hansma, W. C. Kaska and R. M. Laine, J. Am. Chem. Soc. 98, 6064 (1976).

7. Three recent reviews are:
- a) P. K. Hansma, Physics Reports, 30C, 145 (1977).
 - b) W. H. Weinberg, Ann. Div. Phys. Chem. 29, 115 (1978).
 - c) P. K. Hansma and J. Kirtley, Acct. of Chem. Res. 11,
440 (1978).
8. S. Colley and P. K. Hansma, Rev. Sci. Instrum. 48, 1192 (1977).
9. D. J. C. Yates, L. L. Murrell and E. B. Prestridge,
J. Catalysis 57, 41 (1979); J. T. Yates, Jr.,
T. M. Duncan, S. D. Worley and R. W. Vaughn,
J. Chem. Phys. 70, 1219 (1979).
10. R. C. Jaklevic and M. R. Gaerttner, Appl. Phys. Lett. 30,
646 (1977).

Figure Captions

Figure 1. The low energy region of the model rhodium/alumina catalyst spectrum is shown for equal exposures to four mixtures of D_2S and CO. The rhodium is highly dispersed. With exposure to 100% CO one peak forms at 51.6 meV that is due to the absorption of two CO molecules per rhodium atom.⁶ With increasing D_2S exposure, a second peak near 60 meV also forms due to the absorption of only one CO molecule per rhodium atom.⁶

Figure 2. The low energy region of the model rhodium/alumina catalyst spectrum is shown for three different exposures to CO. The rhodium is not highly dispersed, but is in approx. 30 Å diameter particles. For CO exposures of a Langmuir or less, a peak near 58 meV, due to a single CO molecule per rhodium atom dominates.⁶ For higher CO exposures, a peak at 51.6 meV due to two CO molecules per rhodium atom rapidly becomes the most intense.⁶

CO WITH D₂S ON
RHODIUM/ALUMINA

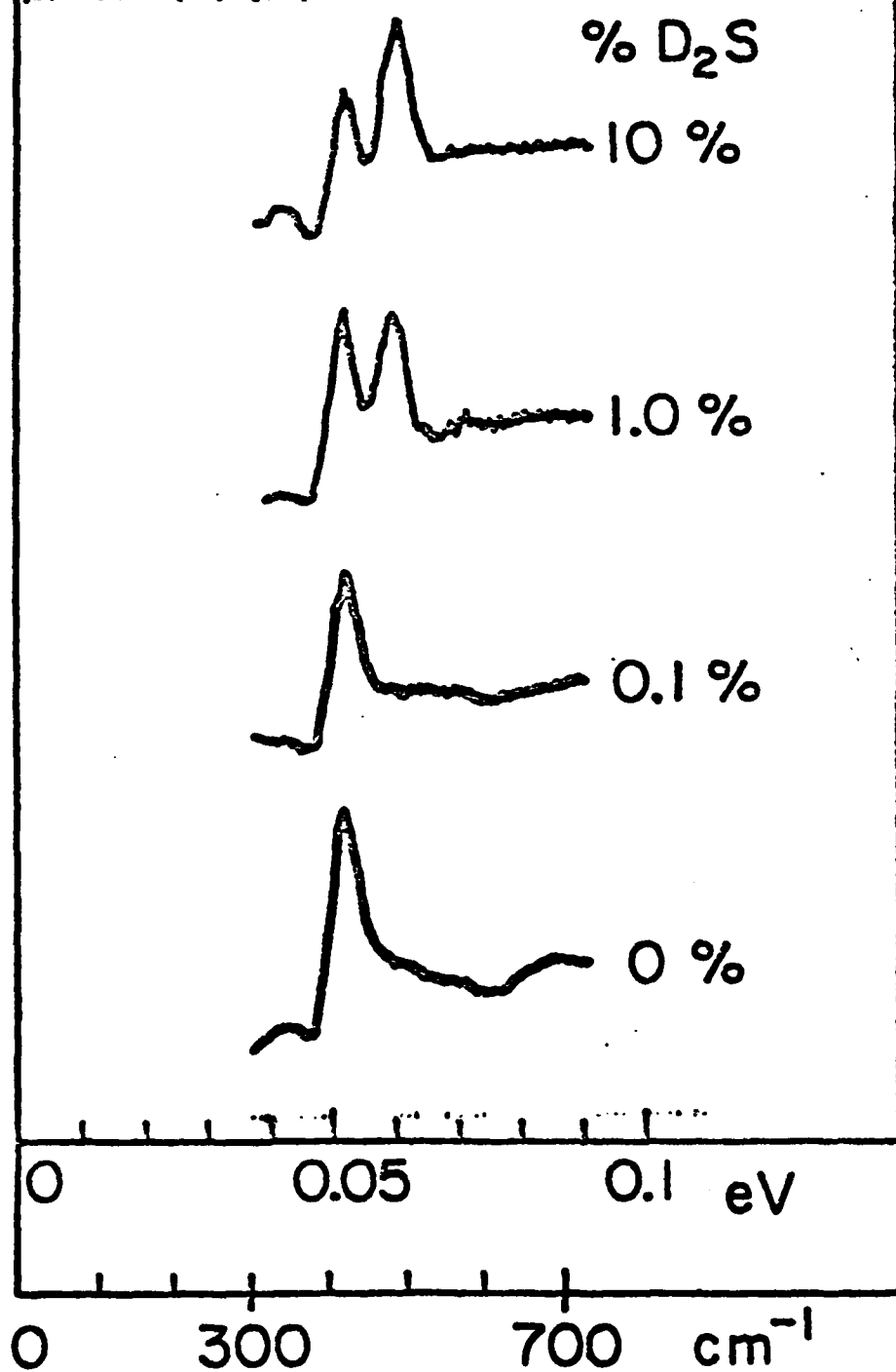


Figure 1

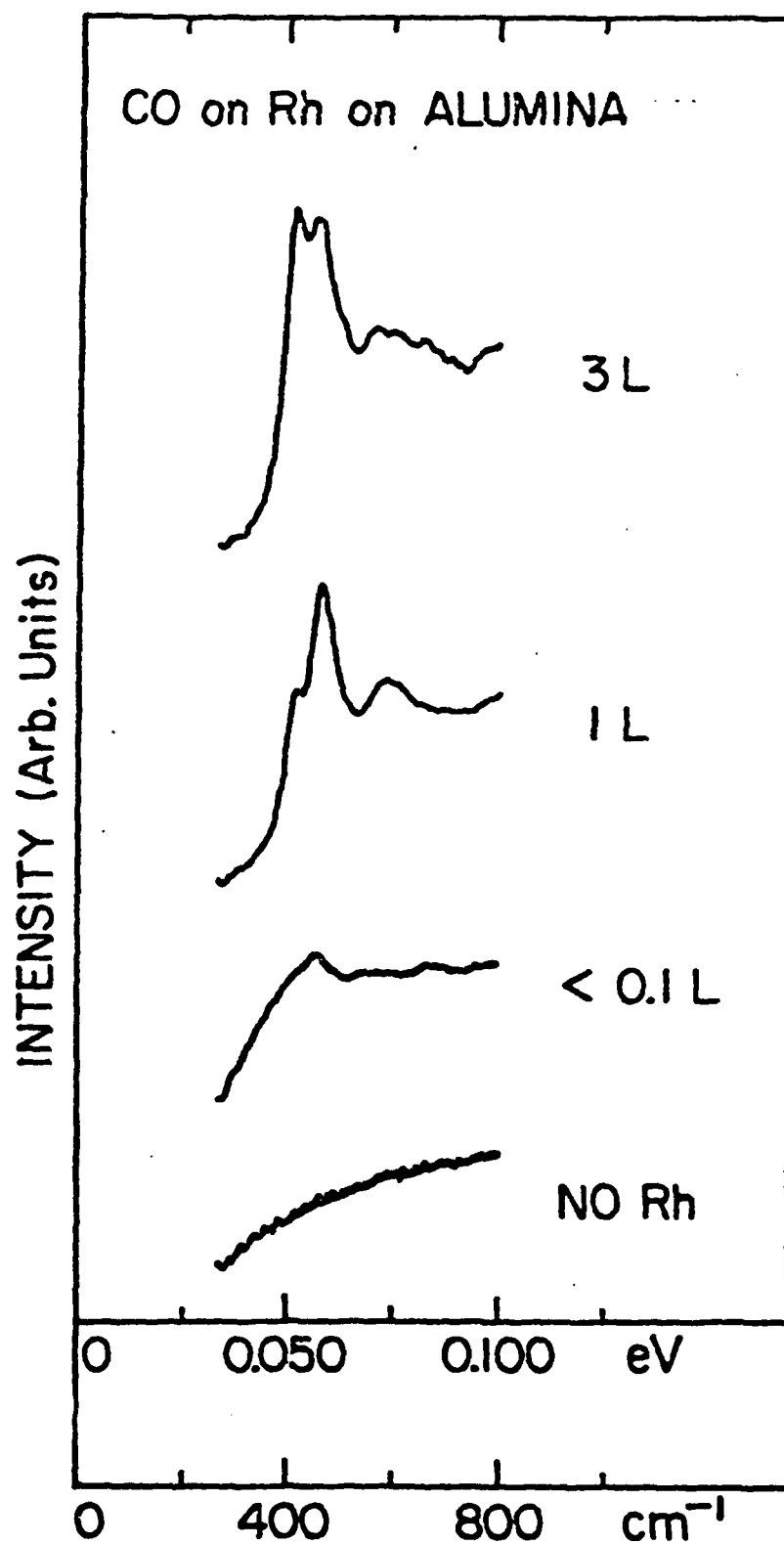


Figure 2